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4.2 For the reaction, 2A + B → A<sub>2</sub>B, the rate = k [A][B]<sup>2</sup> with k = 2.0 × 10<sup>-6</sup> mol<sup>-2</sup> L<sup>2</sup> s<sup>-1</sup>. Calculate the initial rate of the reaction when [A] = 0.1 mol L<sup>-1</sup>, [B] = 0.2 mol L<sup>-1</sup>. Calculate the rate of reaction after [A] is reduced to 0.06 mol L<sup>-1</sup>. Sol. Initial rate of reaction = k [A] [B]<sup>2</sup> = (2.0 × 10<sup>-6</sup> mol<sup>-2</sup> s<sup>-1</sup>) (0.1 mol L<sup>-1</sup>) (0.2 mol L<sup>-1</sup>)<sup>2</sup> = 8 × 10<sup>-9</sup> mol<sup>-1</sup> s<sup>-1</sup>.

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