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the quantum mechanical model where

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electrons do not follow specific paths, but are given probabilities of being found in a particular area-- orbitals, clouds; recognized wave-particle duality
Orbital the area in which an electron has a 90% probability of being found

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Chapter 13 - Electrons in Atoms.

Chapter 13: 1 - 20, 23 - 25, 27, 31, 32, 34 - 38, 41, 45, 47, 48, 52. Section 13.1

-Models of the Atom. Section Review

13.1. 1. List in chronological order, a major contribution of each of these scientists to the understanding of the atom: proposed that all elements are composed of atoms.

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CHEMSITRY NOTES - Chapter 13

Electrons in Atoms Goals : To gain an understanding of : 1. Atoms and their structure. 2. The development of the atomic theory. 3. The quantum mechanical model of the atom. 4.

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Electron configurations NOTES: The different historical models are described as follows:

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quantum. ... the regions around the nucleus within which the electrons have the highest probability of being found.
energy level. a region in space around the nucleus of an atom where an electron is likely to be moving. quantum mechanical model.

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