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## Empirical And Molecular Formula

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EMPIRICAL AND  
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WORKSHEET An oxide  
of chromium is found  
to have the following %  
composition: 68.4 % Cr  
and 31.6 % O.

Determine this

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## Empirical And Molecular Formula

compound's empirical formula. The percent composition of a compound was found to be 63.5 % silver, 8.2 % nitrogen, and 28.3 % oxygen. Determine the compound's empirical formula.

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## Empirical And Molecular Formula

WORK ON A SEPARATE SHEET OF PAPER. Write the empirical formula for the following

- compounds. 1)  $C_6H_6$   
2)  $C_8H_{18}$  3)  $WO_2$  4)  $C_2H_6O_2$  5)  $X_3Y_{13}$  6)

A compound with an empirical formula of  $C_2OH_4$  and a molar mass of 88 grams per mole.

## **Empirical and Molecular Formula Worksheet**

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## Empirical And Molecular Formula

Molecular Formula

Worksheet ANSWER

KEY. Write the  
empirical formula for  
the following

compounds. 1) C<sub>6</sub>H<sub>6</sub>

CH. 6) C<sub>8</sub>H<sub>18</sub> C<sub>4</sub> H<sub>9</sub>

7) WO<sub>2</sub> WO<sub>2</sub> 8)

C<sub>2</sub>H<sub>6</sub>O<sub>2</sub> CH<sub>3</sub> O<sub>9</sub>)

X<sub>39</sub>Y<sub>13</sub> X<sub>3</sub>

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Empirical and  
molecular formula  
worksheet. What is the

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## Empirical And Molecular Formula

empirical formula for a compound that contains 0.063 mol chlorine and 0.22 mol. 10 min maximum page 335 question 52. 52 a compound was found to contain 49.98 g carbon and 10.47 g hydrogen. E te 4i16 2.

## **Empirical And Molecular Formula Worksheet - worksheet**

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WORK ON A SEPARATE SHEET OF PAPER. Write the empirical formula for the following

compounds. 1) C<sub>6</sub> H<sub>6</sub> . 2) C<sub>8</sub> H<sub>18</sub> C<sub>4</sub> H<sub>9</sub> . 3) WO<sub>2</sub> WO<sub>2</sub> . 4) C<sub>2</sub> H<sub>6</sub> O<sub>2</sub> CH<sub>3</sub> O . 5) X<sub>39</sub> Y<sub>13</sub> X<sub>2</sub> Y.

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exam style questions  
like determine the  
empirical formula of a  
compound that  
contains 53.70% iron  
and 46.30% sulfur, ...  
solutions are provided  
at the end of each  
worksheet.

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## **Empirical Formula Worksheets - Answers 8.10 DSoftSchools**

Find the empirical and molecular formulas for this compound. A compound contains 62.0% C, 10.4% H, and 27.5% O. Determine the empirical formula for this compound. After analysis, it was found that the compound's molar mass is 58.1g.

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**MOLECULAR  
FORMULAS  
WORKSHEET**

Percent Composition  
and Molecular Formula  
Worksheet Key 1.

What's the empirical  
formula of a molecule  
containing 65.5%  
carbon, 5.5%

hydrogen, and 29.0%  
oxygen?  $C_3H_3O$

mass = 55 g/mole 2. If  
the molar mass of the  
compound in problem  
1 is 110 grams/mole,  
what's the molecular

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formula? C 6 H 6 O 2 3.  
Worksheet

**Percent Composition  
and Molecular**

**Formula Worksheet**

Molecular Formula  
Worksheet Molecular  
formula - a formula  
showing the types and  
numbers of atoms  
combined in a single  
molecule of a  
molecular compound. It  
is a whole number  
multiple of the  
empirical formula.

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...

What is the empirical formula of rubbing alcohol? Empirical and Molecular Formula Worksheet ANSWER KEY. Write the empirical formula for the following compounds. 1)  $C_6H_6$  .  $C_3H_3$ . 2)  $C_8H_{18}$  .  $C_4H_9$  . 3)  $WO_2$  .  $WO_2$  . 4)  $C_2H_6O_2$  .  $CH_3O$  . 5)  $X_{39}Y_{13}$  .  $X_3Y$  . 6) A

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## Empirical And Molecular Formula

compound with an empirical formula of  $C_2O_4H_4$  and a molar mass of 88 grams per mole.

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Teacher ...**

Worksheet #8

Empirical Formulas 1.

State the empirical  
formula for each of the  
following compounds:

a)  $C_4H_8$ ; b)  $C_2H_6O_2$ ;

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## Empirical And Molecular Formula

c)  $\text{N}_2\text{O}_5$ ; d)  $\text{Ba}_3(\text{PO}_4)_2$ ; e)  $\text{Te}_4\text{I}_{16}$  2. What is the empirical formula for a compound that contains 0.063 mol chlorine and 0.22 mol

### **Worksheet #8**

### **Empirical Formulas**

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2. Determine the  
empirical formula from  
the percent  
composition for each of  
the following: a) 92.24  
% C; 7.76 % H. b)  
36.48 % Na; 25.44 % S;

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